LOHS • AP Chemistry • Schrempp

Name _____
Period ___ Date ___/___/___

4.3 • Reaction in Aqueous Solution

ANSWERS to NAMING ACIDS

"ate" becomes "_____ic acid"
"ite" becomes "_____ous acid"
"ide" becomes "hydro_____ic acid

in sulfur compounds, add "ur"
in phosphorus compounds, add "or"

bromate	BrO_3^-	$HBrO_3$	bromic acid	perchlorate	ClO_4^-	$HClO_4$	perchloric acid
periodate	IO_4^-	HIO_4^-	periodic acid	bisulfate*	HSO_4^-	H_2SO_4	sulfuric acid
carbonate	CO_3^{2-}	H_2CO_3	carbonic acid	hypoiodite	IO^-	HIO	hypoiodous acid
peroxide*	O_2^{2-}	H_2O_2	hydrogen peroxide	bicarbonate*	HCO_3^-	H_2CO_3	carbonic acid
chloride	Cl^-	HCl	hydrochloric acid	sulfate	SO_4^{2-}	H_2SO_4	sulfuric acid
chlorite	ClO_2^-	$HClO_2$	chlorous acid	iodite	IO_2^-	HIO_2	iodous acid
thiosulfate	$S_2O_3^{2-}$	$H_2S_2O_3$	thiosulfuric acid	acetate	$C_2H_3O_2^-$	$HC_2H_3O_2$	acetic acid
sulfide	S^{2-}	H_2S	hydrosulfuric acid	iodide	I^-	HI	hydroiodic acid
dichromate	$Cr_2O_7^{2-}$	$H_2Cr_2O_7$	dichromic acid	bromide	Br^-	HBr	hydrobromic acid
hypobromite	BrO^-	$HBrO_2$	hypobromous acid	hydroxide*	OH^-	HOH	water
sulfite	SO_3^{2-}	H_2SO_3	sulfurous acid	phosphate	PO_4^{3-}	H_3PO_4	phosphoric acid
chromate	CrO_4^{2-}	H_2CrO_4	chromic acid	hypochlorite	ClO^-	HClO	hypochlorous acid
permanganate	MnO_4^-	$HMnO_4$	permanganic acid	phosphite	PO_3^{3-}	H_3PO_3	phosphorous acid
iodate	IO_3^-	HIO_3^-	iodic acid	oxide*	O^{2-}	H_2O	water
perbromate	BrO_4^-	$HBrO_4$	perbromic acid	fluoride	F^-	HF	hydrofluoric acid
cyanide	CN^-	HCN	hydrocyanic acid	thiocyanate	SCN^-	HSCN	thiocyanic acid
chlorate	ClO_3^-	$HClO_3$	chloric acid	bromite	BrO_2^-	$HBrO_2$	bromous acid
nitrate	NO_3^-	HNO_3	nitric acid	nitrite	NO_2^-	HNO_2	nitrous acid

* = be careful